

[OSSO]-Type Iron(III) Complexes for the Low-Pressure Reaction of Carbon Dioxide with Epoxides: Catalytic Activity, Reaction Kinetics and Computational Study

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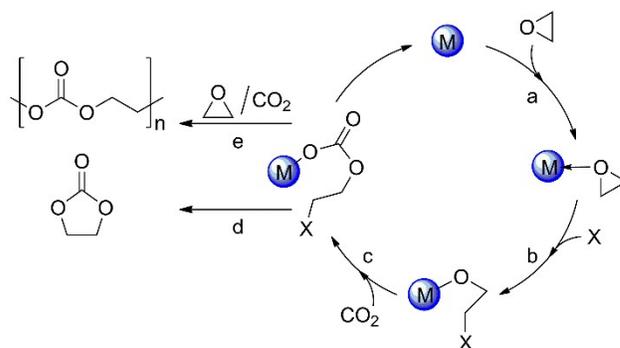
ABSTRACT: The selective conversion of variously substituted epoxides into the corresponding cyclic carbonates (COCs) under mild reaction conditions was achieved with mononuclear Fe(III) complexes bearing bis-thioether-di-phenolate [OSSO]-type ligands, in combination with tetrabutylammonium bromide (TBAB). For example, propylene carbonate was obtained in 6 h at 35 °C with a conversion of 65% (TOF = 108 h⁻¹), from propylene oxide and 1 bar of CO₂ pressure, using 0.1 mol% of the Fe(III) complex and 0.5 mol% of TBAB. Product divergence is observed only for cyclohexene oxide toward the exclusive formation of the aliphatic polycarbonate (APC) (TOF = 165 h⁻¹ at 80 °C and 1 bar of CO₂ pressure, using 0.1 mol% of the Fe(III) complex and 0.1 mol% of TBAB). Kinetic investigations indicated reaction orders of two and one, with respect to the Fe(III) complex, for the production of propylene carbonate and the poly(cyclohexene carbonate), respectively. The enthalpy and entropy of activation were determined using the Eyring equation (for propylene carbonate: $\Delta H^\ddagger = 8.4 \pm 0.7$ kcal/mol and $\Delta S^\ddagger = -33 \pm 3$ cal/(mol · K); for poly(cyclohexene carbonate): $\Delta H^\ddagger = 11.9 \pm 0.3$ kcal/mol and $\Delta S^\ddagger = -36 \pm 2.2$ cal/(mol · K)). Supported by density functional theory based investigations, we propose a mechanistic scenario in which the rate limiting step is the bimetallic ring opening of the epoxide, in the case of propylene carbonate, and the monometallic insertion of the epoxide in the growing polymer chain, in the case of poly(cyclohexene carbonate).

INTRODUCTION

The last two decades have witnessed growing interest in the capture and utilization of carbon dioxide (CO₂).¹ This strategy, the so-called carbon dioxide utilization (CDU), offers a conceptually simple way to reduce anthropogenic CO₂ emissions, with the advantage of using an abundant non-toxic C1 chemical feedstock for the production of added-value products, such as fuels, chemicals and materials.² However, implementing processes based on the CO₂ chemistry is hampered by the thermodynamic stability and the kinetic inertness of this molecule.³ Two of the principal strategies to overcome the thermodynamic stability of CO₂ are the electrochemical reduction of CO₂, where the applied potential provides the thermodynamic driving force,⁴ or targeting reactions where CO₂ is coupled with a thermodynamically unstable molecule, which provides the thermodynamic driving force towards value added chemicals.⁵ A representative example of this concept is the coupling of CO₂ with epoxides to give cyclic organic carbonates (COCs),⁶ or aliphatic polycarbonates (APCs).⁷ COCs have many applications as highly boiling polar solvents,⁸ electrolytes in lithium-ion batteries,⁹ and chemical intermediates in the synthesis of other industrially relevant products.¹⁰ On the other hand APCs have several applications as materials,¹¹ and can also be used to store CO₂ for a longer range of time with

respect to small molecules, due to the long lasting life of these polymers.¹²

Scheme 1. General mechanism for the coupling of CO₂ and epoxides^a



^aAbbreviations: M = metal complex; X = nucleophile.

The kinetic inertness of CO₂ can be overcome by developing effective catalytic systems. Those for the synthesis of COCs and APCs are generally based on the combination of a transition metal complex with a nucleophile (quaternary ammonium salts and phosphonium salts) as co-catalyst, with several met-

als known to perform effectively (e.g. Al, Zn, Co, Cr).^{6e,7f} It is widely accepted that these reactions proceed through a mechanism involving the following fundamental steps (Scheme 1):^{6b} (a) the epoxide ring coordinates to the Lewis acidic metal center; (b) the activated epoxide reacts with the nucleophilic co-catalyst to generate a metal-alcoholate species; (c) the insertion of CO₂ into the metal-oxygen bond leads to the formation of an hemi-carbonate intermediate. The latter may undergo two different reactions: (d) ring closure, with formation of the cyclic carbonate or; (e) the alternate insertion of the epoxide and CO₂, affording polycarbonate architectures.

In spite of several studies aimed to clarify the mechanistic aspects of the CO₂/epoxide coupling reaction a widely accepted mechanistic scenario, including the comprehension of the factors governing the selectivity toward the cyclic or polymeric product, are still lacking. Nevertheless, in many cases the use of a bimetallic catalytic system with two proximal metal centers seems to have a fundamental role in both the cyclization and co-polymerization reactions.¹³ In the first case, North and Pasquale reported a bimetallic aluminum(salen) complex for the synthesis of styrene carbonate, in which one metal center activates the epoxide, and the other one promotes the CO₂ insertion into the metal-alkoxide bond.^{13a} In the second case, Williams et al. described the use of a bimetallic zinc complex for the production of poly(cyclohexene carbonate), claiming that a bimetallic mechanism is crucial for both the initiation and propagation steps.^{13b} Specifically, during the initiation step one metal center activates the epoxide and the second one furnishes the nucleophile for the ring-opening, while in the propagation step one metal activates the epoxide and the second one, bearing the growing polycarbonate chain, activates the carbonate end-group for the nucleophilic attack.^{13c} The presence of a dinuclear catalytic system is also reported to prevent the back-biting reaction responsible for the formation of mixtures of cyclic and polymeric products.^{7c} The occurrence of a bimetallic mechanism, was demonstrated even in the presence of mononuclear complexes, as reported by Coates,¹⁴ and Rieger,¹⁵ for the production of poly(propylene carbonate) and propylene carbonate, respectively. Recently, we have reported on a family of dinuclear Fe(III) complexes coordinated by bis-thioether-tri-phenolate ligands that, in combination with tetrabutylammonium bromide (TBAB), have shown to be highly active in the synthesis of various COCs from terminal and internal epoxides.¹⁶ Intriguingly, mechanistic studies through DFT based calculations have not only evidenced the crucial role of the hemilabile Fe-S bond in the catalytic activity, but also that only one metal center is active in the catalytic cycle.¹⁷

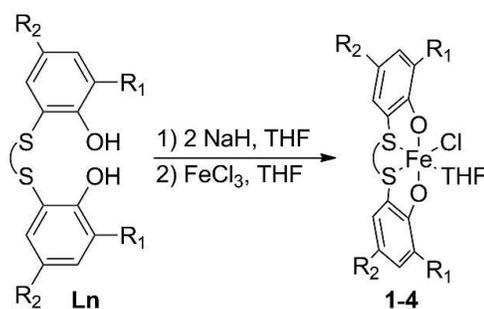
Inspired from these findings, and encouraged by the fact that different examples of highly active catalysts in both the cyclization and co-polymerization reactions have been reported for which a monometallic mechanism has been proposed,¹⁸ our attention has thus moved to the synthesis of mononuclear iron(III) complexes containing sulfur atoms in the ligand backbone. In particular, in this work we report on the synthesis and complete characterization of a new family of [OSSO]-type iron(III) complexes and their use, when combined with suitable halide nucleophiles, as highly efficient binary catalysts for COCs formation for a wide range of substrates under very mild reaction conditions, recording the highest turnover frequency (TOF) values ever reported under 1 bar of CO₂ pressure. Furthermore, these catalysts also promote the efficient co-polymerization of cyclohexene oxide with CO₂ giving the cor-

responding polycarbonates with highest activity and selectivity so far reported for iron based catalysts. A complete kinetic investigation was conducted for both the cyclization and co-polymerization reactions and a mechanistic scenario is proposed, combining experimental evidences with DFT based calculations.

RESULTS AND DISCUSSION

Synthesis and Characterization of the [OSSO]-Fe(III) complexes 1-4. The bis-thioether-diphenolate [OSSO]-type pro-ligand **L1**, with bulky trityl substituents in the R₁ positions, and pro-ligands **L2-L4**, were synthesized according to the reported procedures (see SI).¹⁹ The mononuclear [OSSO]-Fe complexes **1-4** were obtained by reaction of FeCl₃ with the sodium salt of the corresponding pro-ligand in THF solution; all the products were recovered in high yield as deep-blue powders (Scheme 2).

Scheme 2. Synthesis of the Iron(III) complexes 1-4



	R ₁	R ₂
1	(CH ₂) ₂	CPh ₃
2	(CH ₂) ₂	C(CH ₃) ₂ Ph
3	(CH ₂) ₂	CH ₃
4	<i>c</i> -Hex	C(CH ₃) ₃

Single crystals of the complex **2** were grown from a saturated acetonitrile solution, and the crystal structure is shown in Figure 1.

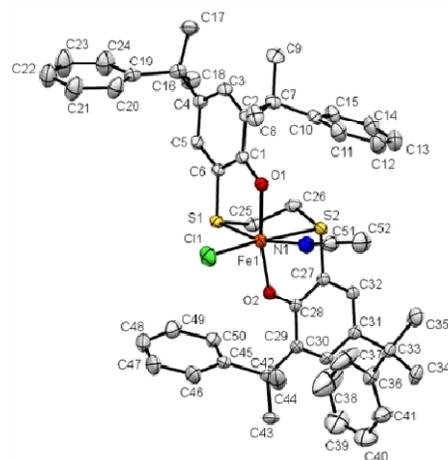


Figure 1. Molecular structure of compound **2**. Hydrogen atoms have been omitted for the sake of clarity (thermal ellipsoids are drawn at 30% of the probability level). Selected bond lengths (Å) and angles (°): Fe₁-O₁ 1.908(2), Fe₁-O₂ 1.900(2), Fe₁-S₁ 2.5642(8), Fe₁-S₂ 2.6936(9), Fe₁-Cl₁ 2.2497(9), Fe₁-N₁ 2.124(2), S₁-Fe₁-S₂ 83.20(2).

The Fe atom has a distorted octahedral geometry being coordinated by one tetradentate [OSSO] ligand adopting a *fac-fac* wrapping coordination mode, one chloride and one N atom from an acetonitrile ligand. The Fe atom is involved in a puckered five-membered ring (Fe₁-S₁-C₂₅-C₂₆-S₂) with a S₁-Fe₁-S₂ bite angle of 83.20(2)°, smaller than that found for example in the complex [Fe(CO)₂(tbu₄-dttdd)].²⁰ The elemental analysis were in agreement with the formation of the desired products, and the disappearance of the hydroxyl group vibration at about 3300 cm⁻¹ is evident in the FT-IR spectra of the complexes (see SI). The formation of a monomeric structure was confirmed by the ESI FT-ICR MS analysis from CH₂Cl₂/MeOH solutions (see SI). In addition the effective magnetic moments (μ_{eff}), determined by means of the Evans method in deuterated toluene in the range from 30 to 90 °C, are comparable with the calculated value for one isolated high spin (HS) Fe(III) center (see SI).

Cycloaddition of CO₂ to epoxides catalyze by 1-4. The cycloaddition of carbon dioxide to 1,2-epoxyhexane (**5a**) in the presence of **1-4** was investigated under very mild reaction conditions, i.e. 35 °C, 1 bar of CO₂ pressure and 0.1 mol% of complex with respect to the epoxide. In the absence of co-catalyst, all the epoxide remained unreacted after 24 hours. One equivalent of TBAB activates the complexes, with the selective conversion of **5a** to the corresponding cyclic carbonate 4-butyl-1,3-dioxolan-2-one (**6a**) (entries 1-4, Table S2). Under these conditions the catalyst **4** resulted the most efficient with a TOF of 13.5 h⁻¹. The variance of the catalytic activity of **4** with respect to the temperature, in the range from 20 to 70 °C, was also monitored (entries 4-7, Table S2). One can observe that the activity increase linearly until 50 °C, while a deviation from linearity was observed at 70 °C. With the aim to find the best catalyst/co-catalyst combination, the screening of the following co-catalyst, in the presence of **4** under the same reaction conditions, was also performed: dimethylaminopyridine (DMAP), bis(triphenylphosphine)iminium chloride ([PPN]Cl), tetrabutylammonium azide (TBAN₃), tetrabutylammonium chloride (TABC) and tetrabutylammonium iodide (TBAI) (entries 8-12, Table S2).

Table 1. 5a/CO₂ coupling promoted by 1-4^a

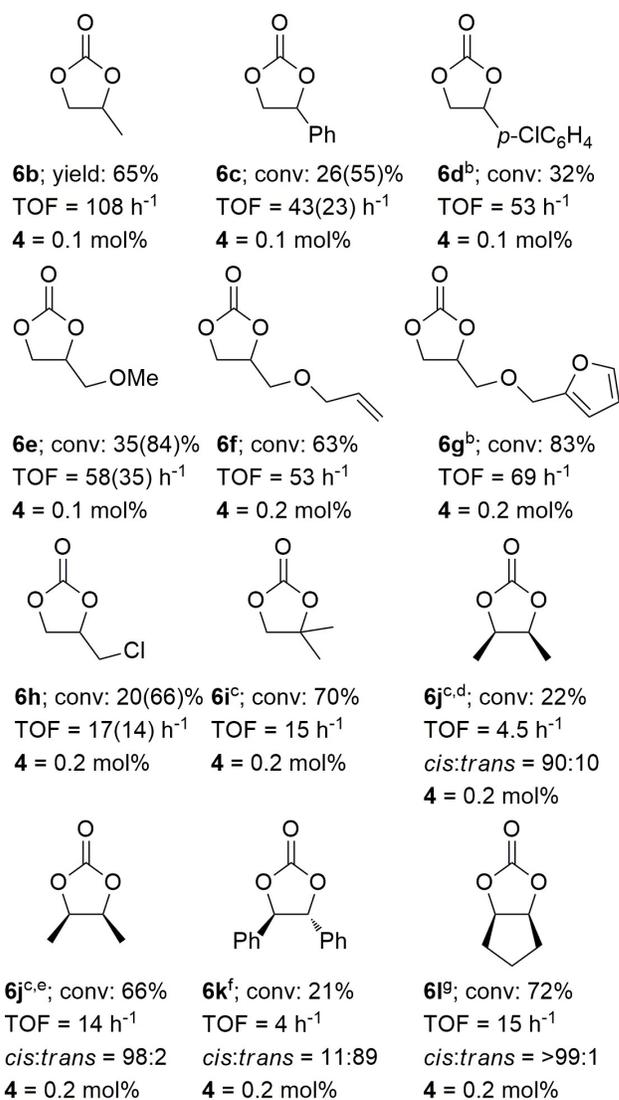
Entry	Catalyst	Conversion (%) ^b	TOF (h ⁻¹) ^c
1	1	33	55
2	2	27	45
3	3	32	53
4	4	38	63
5 ^d	-	2	-

^aReaction conditions: **5a** = 4.15 · 10⁻² mol; Catalyst = 4.15 · 10⁻⁵ mol (0.1 mol%); TBAB = 2.07 · 10⁻⁴ mol (0.5 mol%); T = 35 °C; P(CO₂) = 1 bar; reaction time = 6 h, neat. ^bDetermined by ¹H NMR using mesitylene as internal standard. The selectivity toward the formation of **6a** was always found to be >99%. ^cTurnover frequency (mol_{6a} · mol_{Catalyst}⁻¹ · reaction time⁻¹). ^dControl experiment in the absence of the iron catalyst.

With the exception of DMAP, which poorly activate the complex, the efficiency of the catalytic system does not vary remarkably by changing the nature of the co-catalyst. Notably, increasing the TBAB amount from 0.1 to 0.5 mol% positively affects the catalytic activity, enhancing the conversion from 32 to 89% (entries 4 and 14, Table S2), further increase of the to 1.0 mol% slightly affects the progress of the reaction, reaching a conversion of 93% (entry 15, Table S2).

The catalytic performances of **1-4** were compared determining the conversion of **5a** to **6a** after 6 hours (Table 1). All the complexes showed high activity, being catalyst **4** the most efficient with a TOF equal to 63 h⁻¹.

Scheme 3. Scope in epoxides^a



^aReaction conditions: epoxide = 4.15 · 10⁻² mol; T = 35 °C; P(CO₂) = 1 bar; reaction time = 6 h, neat. Conversion determined by ¹H NMR using mesitylene as internal standard. The selectivity toward the formation of cyclic carbonate was always found >99%. Stereoselectivity determined by ¹H NMR. Values in brackets obtained at 24 h. ^bEpoxide = 1.5 g. ^cEpoxide = 1.0 mL; T = 50 °C; P(CO₂) = 10 bar; reaction time = 24 h. ^d[TBAB]/[**4**] = 1/1. ^e[TBAB]/[**4**] = 5/1. ^fEpoxide = 1.0 g; T = 90 °C; P(CO₂) = 10 bar; reaction time = 24 h. ^gT = 50 °C; P(CO₂) = 10 bar; reaction time = 24 h.

Complex **4** was then tested, in combination with TBAB, for the conversion of a group of variously substituted epoxides in the corresponding COCs (Scheme 3). The reaction of the less encumbered propylene oxide **5b** proceeds faster, with a TOF of 108 h⁻¹. To the best of our knowledge, this is the highest activity reported so far, for a homogeneous catalyst, for the formation of **6b** from **5b** under 1 bar of CO₂ pressure.^{6c} A small attenuation in the catalytic activity was observed in the case of the phenyl substituted substrates **5c** and **5d**. A comparable catalytic activity was observed for the methoxy containing epoxide **5e**. For the production of allyloxy and furfuryl substituted carbonates **6f** and **6g**, the use of 0.2 mol% of **4** was necessary. The formation of **6h** proceeds slower, as reported for other catalytic systems.²⁶

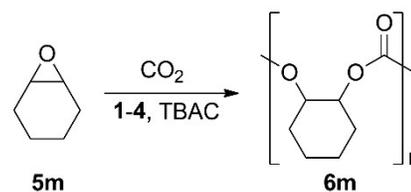
The synthesis of di-substituted carbonates **6i-k**, known to be more challenging to obtain, was successfully accomplished at 50 °C and 10 bar of CO₂ pressure, with good activities. In particular, **6j** was obtained with a high degree of stereoretention with respect to the starting epoxide, even when an equimolar mixture of **4** and TBAB was used. The formation of the bulky carbonate **6k** requires higher temperatures to melt the epoxide, and also in this case the product was obtained with a high degree of stereoretention. The reaction of cyclopentene oxide (**5l**) proceeds as well, with the selective formation of the corresponding cyclic carbonate **6l**. Intriguingly, the reaction of cyclohexene oxide (**5m**), under the same reaction conditions, led to the exclusive formation of almost perfect poly(cyclohexene carbonate) (**6m**) (entry 1 table S4). Such divergence in the selectivity was often reported for this substrate,^{26,18d} but the examples of iron based catalytic systems capable of co-polymerize **5m** with CO₂ are rather rare.²⁷

Co-polymerization of 5m and CO₂ catalyzed by 1-4. In order to get a deeper insight into the behavior of this new family of catalysts, the co-polymerization reaction of CO₂ with **5m** in the presence of **1-4** at 80 °C and 10 bar of CO₂ pressure was performed, using 0.1 mol% of complex and 0.1 mol% of TBAB (entries 2-5, Table S3). Notably in this case the nature of the substituents on the [OSSO] ligands has a more pronounced effect on the catalytic activity with respect to the cycloaddition reaction. Indeed, the methyl substituted complex **3** only leads to the formation of traces of polymer after 2 hours (entry 4, table S3), while a conversion of 40% was obtained in the presence of the more sterically encumbered cumyl substituted complex **2** (entry 3, table S3), with a TOF of 200 h⁻¹. A further increase of the steric demand introducing trityl groups in the R₁ positions, or reducing the flexibility of the ligand, changing the ethyl bridge with a cyclohexyl moiety, negatively affects the catalytic performance respect to catalyst **2** (entries 2 and 5, table S3). The effect of co-catalyst type and loading on the catalytic activity of complex **2** was also investigated (entries 6-18, table S3). It is worth noting that in the case of the co-polymerization reaction the catalyst, although with lower activity, gives the desired product also in the absence of any co-catalyst with a lower content of carbonate linkages (entry 7, table S3). Moreover TBAC vis-à-vis TBAB was found to be the most effective co-catalyst, reaching 34% of conversion in 1 hour and a TOF equal to 340 h⁻¹ (entry 16, table S3). Such difference is often observed for the co-polymerization reaction, and is commonly attributed to the different behavior of the chloride anion respect to the bromide as a leaving group.^{6b} Halving the amount of TBAC with respect to **2** halves the catalytic activity but it does not affect the polymer composition (entry 15, table S3), while using two equivalents of TBAC

with respect to iron, resulted only in a moderate increase of the catalytic activity, reaching a TOF equal to 400 h⁻¹. The variation of the catalytic activity of **2** with respect to the temperature, in the range from 40 to 100 °C, was also monitored (entries 19-21, Table S3). When the temperature was lower than 60 °C, the conversion was negligible. Raising the temperature to 100 °C increase the catalytic activity of about the 20%. The effect of the carbon dioxide pressure on the co-polymerization process was investigated. When the reaction was performed at 5 bar, the catalytic activity was slightly lower than that recorded at 10 bar. Remarkably, good catalytic activity was observed even at 1 bar of CO₂ pressure (TOF = 165 h⁻¹, entry 22, Table S3), without variations in the polymer composition. Interestingly the TOF observed at 20 bar was similar to that of the experiment conducted at 1 bar (entry 24, Table S3), however significantly lower molecular weight were obtained in the first case.

At this point, the catalytic performances of **1-4** were compared determining the conversion of **5m** in **6m** after 1 hour, at 80 °C, 10 bar of CO₂ pressure, 0.1 mol% of iron and 0.1 mol% of TBAC (entries 1-4, Table 2). Only for complex **3**, the reaction time was fixed at 6 hours in order to obtain a suitable polymer sample for further analysis and to verify the selectivity of the reaction (entry 3, Table 2). Under these optimized reaction conditions, the trend of the catalytic activities observed by using TBAB as co-catalyst was confirmed, with complex **2** being the most effective catalyst of the series (entry 2, Table 2).

Table 2. 5m/CO₂ co-polymerization using 1-4^a



Entry	Cat.	Yield (%) ^b	TOF ^c	Mn (kDa) ^d	PDI ^d	Mn1/Mn2 (%) ^d
1	1	14	140	9.8	1.01	7
				4.6	1.04	93
2	2	34	340	23.2	1.01	30
				11.0	1.03	70
3 ^e	3	25	42	20.0	1.01	15
				10.4	1.02	85
4	4	17	170	17.8	1.01	24
				8.3	1.04	76
5 ^f	2	29	290	14.3	1.01	28
				6.7	1.03	72
6 ^g	2	40	400	27.8	1.01	7
				12.8	1.06	93

^aReaction conditions: **5m** = 5.93 · 10⁻² mol; catalyst = 0.1 mol%; TBAC = 0.1 mol%; T = 80 °C; P(CO₂) = 10 bar; reaction time = 1 h; neat. ^bGravimetric. Carbonate linkages always >99%. ^cTurnover Frequency (yield · 0.01 · mol_{Catalyst}⁻¹ · reaction time⁻¹). ^dDetermined by Gel Permeation Chromatography at 30 °C in THF respect to polystyrene standards. ^ereaction time = 6 h. ^fH₂O = 0.1 mol%. ^g**5m** was distilled twice over CaH₂ prior to use.

In all cases a bimodal distribution of the molecular weights was obtained. Such behavior is commonly observed for this

kind of polymerization, and was typically attributed to the presence of traces of water in the reaction system.^{13b,27c} To verify this hypothesis the co-polymerization of **5m** with CO₂ in the presence of **2** was repeated under two different reaction conditions: in the first case one equivalent of water respect to iron was added to the reaction mixture, in the second case the amount of water in the system was minimized by distilling the epoxide twice over CaH₂ (entry 5 and 6, Table 2). In the presence of water the catalytic system is still capable to promote the co-polymerization, however lower molecular weights were obtained. On the contrary, minimizing the water content in the system reflects favorably on both catalytic activity and molecular weights distribution. Indeed the TOF raises from 340 to 400 h⁻¹ and an almost unimodal and narrow distribution was obtained. To the best of our knowledge, this is the highest activity reported so far for an iron based catalyst for the co-polymerization of CO₂ and **5m**. This is also beneficial for the thermal properties of the final product. In effect the glass transition temperature (T_g) increases from 91 °C to 118 °C as a consequence of the increased molecular weight of the polycarbonate (Figure S26).

To better understand the reaction mechanism, a detailed kinetic investigation was performed for the formation of **6b** using the system **4**/TBAB and for the **6m** formation using the system **2**/TBAC, by means of online attenuated total reflection infrared (ATR-IR) measurements.

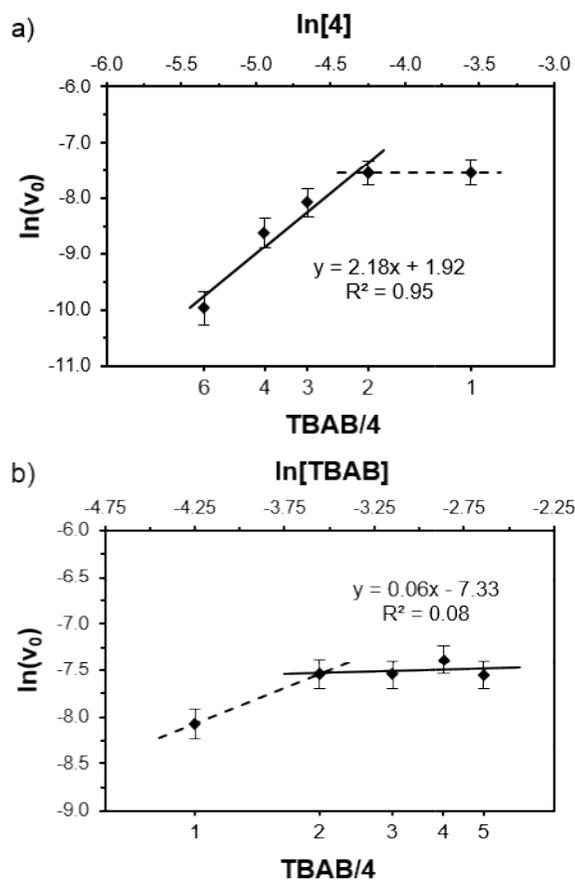


Figure 2. Plot of $\ln(v_0)$ for the formation of **6b** versus: (a) $\ln[4]$, (b) $\ln[TBAB]$. The TBAB/4 molar ratio is reported on the bottom axes.

Kinetic study for **6b formation.** The formation of **6b** in the presence of **4**/TBAB as the catalytic system was investigated, using a **5b**/CH₂Cl₂ solution as the reaction medium to guarantee the homogeneity of the system during the experiments. The general kinetic equation for the reaction of CO₂ with **5b** can be written as follow:

$$v = k \cdot [5b]^a \cdot [CO_2]^b \cdot [4]^c \cdot [TBAB]^d \quad (\text{eq. 1})$$

At first, the reaction order with respect to **4** was determined by performing a series of experiments varying the concentration of **4** in the range from 4.8 to 28.6 mM, keeping the concentration of TBAB constant at 28.6 mM. The double logarithmic plot of the initial rate (v_0) versus the concentration of the complex is shown in Figure 2a. It emerges that the rate of the reaction follows a second-order dependence in **4**. However the dependence on the catalyst concentration is valid only in the presence of an excess of the co-catalyst, namely when the TBAB/4 molar ratio is equal or major than 2. This finding is in accordance with the observation that the presence of the co-catalyst is essential for the expression of the catalytic activity.

An analogous set of experiments was performed to study the dependence with respect to TBAB, varying the concentration of TBAB in the range from 14.3 to 71.5 mM, keeping the concentration of **4** constant at 14.3 mM. The double logarithmic plot of the initial rate versus the concentration of the co-catalyst is shown in Figure 2b. Compared with **4**, a specular behavior was observed in this case. Indeed the reaction rate increases until the TBAB is in excess with respect to iron, while a zero-order dependence was observed when the TBAB/4 molar ratio is equal or higher than 2. The results described above suggest that complex **4** reacts with TBAB forming a new species that is involved in the reaction mechanism. A similar behavior was recently described by North et al. in the case of a chromium(III) salophen system.^{18d}

In that case the reaction order with respect to the concentration of the chromium complex was found to be zero when the concentration of Cr(III) was higher than that of TBAB, and one when it was lower. Such behavior was explained showing that the native complex reacts with TBAB forming a new species $[Cr(\text{salophen})Br_2]$, that was claimed to be the real catalyst.

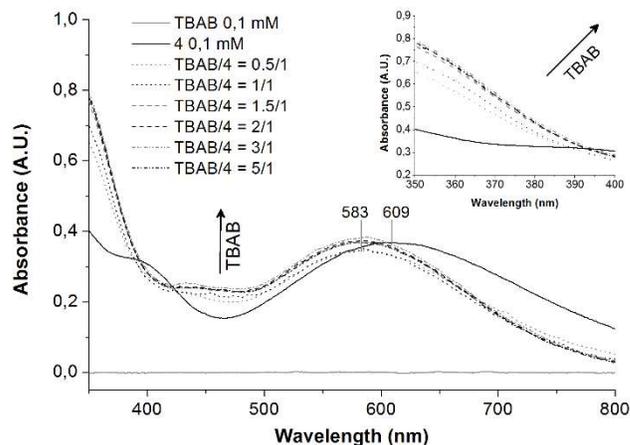


Figure 3. UV-Vis spectra of complex **4** in CH₂Cl₂ solution before and after the addition of TBAB (from 0.5 to 5 equivalents with respect to Fe(III)). The TBAB spectrum is reported on the bottom for comparison. Inset: magnification of the spectra from 350 to 400 nm.

To verify the occurrence of a similar reaction, the UV-Vis spectra of CH_2Cl_2 solutions of **4** in the presence of different amounts of TBAB have been recorded (Figure 3).

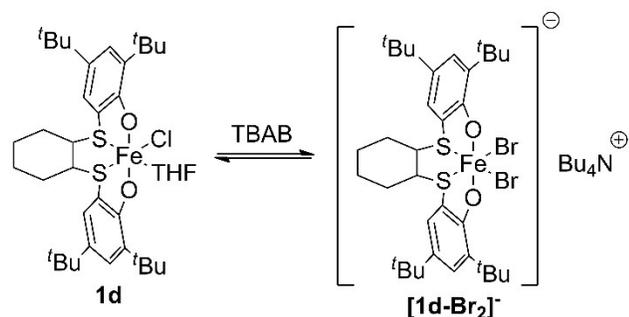


Figure 4. Formation of the anionic species $[\mathbf{4}\text{-Br}_2]^-$ in the presence of TBAB.

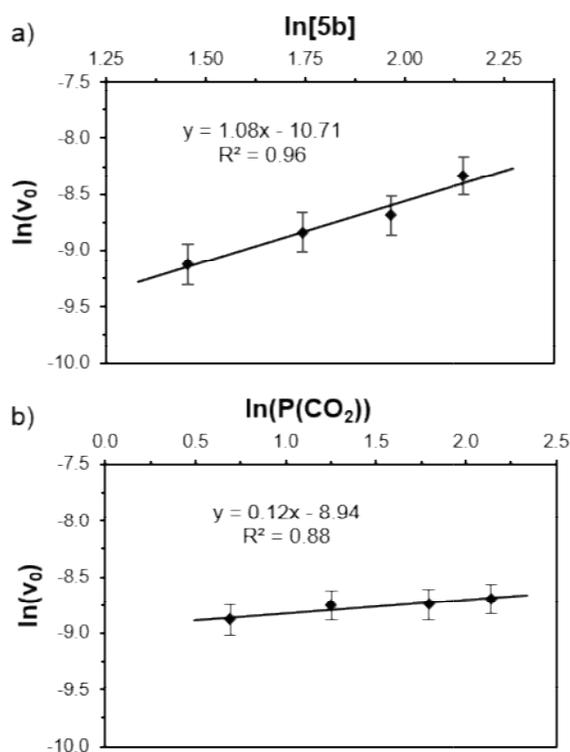


Figure 5. Plot of $\ln(v_0)$ for the formation of **6b** versus: (a) $\ln[\mathbf{5b}]$, (b) $\ln(P(\text{CO}_2))$.

After the addition of only 0.5 equivalent of TBAB, the spectrum profile completely changes. For example the maximum of absorbance at 610 nm undergoes a blue-shift to 583 nm. It is worth noting that the absorbance intensity raises by adding more TBAB, until the $\text{TBAB}/\mathbf{4}$ molar ratio is equal to 2 (see the inset of Figure 3). At this value the intensity does not change anymore, even in the presence of 5 equivalents of TBAB. From this results it appears that the reaction seems to proceed until 2 equivalent of TBAB with respect to iron are added to the solution. This can be explained with the formation a new iron containing anionic species $[\mathbf{4}\text{-Br}_2]^-$ in which the chlorine atom is substituted with the bromine coming from TBAB, and consequently further reaction is not possible (Figure 4).

The reaction order with respect to **5b** was determined performing a series of experiments with different concentration of the

epoxide. The double logarithmic plot of v_0 with respect of the concentration of **5b** gives a first-order dependence (Figure 5a). With constant epoxide concentration, the CO_2 pressure was changed in the range from 2 to 8.5 bar, and the reaction order with respect to CO_2 was found to be zero (Figure 5b).

On the basis of the Eyring equation, the thermodynamic parameters for the reaction can be determined plotting the $\ln(k/T)$ versus $1/T$. The reaction of **5b** with CO_2 was conducted, under the same reaction conditions, at different temperature from 30 to 70 °C, and the Eyring plot is shown in Figure S35. We found the following results: $\Delta H^\ddagger = 8.4 \pm 0.7$ kcal/mol, $\Delta S^\ddagger = -33$ cal/(mol · K) and $\Delta G^\ddagger_{323} = 19.0 \pm 1.7$ kcal/mol.

Finally, in the presence of an excess of the co-catalyst, the kinetic equation for the formation of **6b** can be written as follow:

$$v = k \cdot [\mathbf{5b}] \cdot [\mathbf{4}]^2 \quad (\text{eq. 2})$$

From all the data described above, it is reasonable to conclude that the ring opening of the epoxy ring is the rate limiting step of the reaction and that, in spite of the mononuclear nature of the $[\text{OSSO}]\text{-Fe(III)}$ complex, the second order dependence with respect to the concentration of **4** shows that two metal centers are involved in this step.

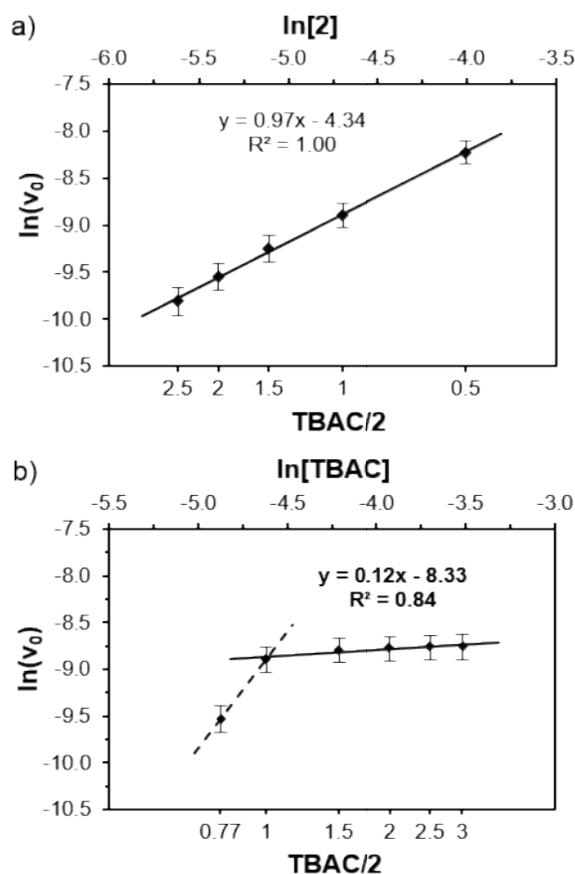


Figure 6. Plot of $\ln(v_0)$ for the formation of **6m** versus: (a) $\ln[\mathbf{2}]$, (b) $\ln[\text{TBAB}]$. The $\text{TBAC}/\mathbf{2}$ molar ratio is reported on the bottom axes.

Kinetic study for 6m formation. The formation of **6m** in the presence of $\mathbf{2}/\text{TBAC}$ as the catalytic system was investigated, using a **5m**/toluene solution as the reaction medium to guarantee the homogeneity of the system during the experiments. The general kinetic equation for the chain growth of **6m** in the copolymerization of CO_2 with **5m** can be written as follow:

$$v = k \cdot [5\mathbf{m}]^a \cdot [\text{CO}_2]^b \cdot [2]^c \cdot [\text{TBAC}]^d \quad (\text{eq. 3})$$

At first, the reaction order with respect to **2** was determined by performing a series of experiments varying the concentration of **2** in the range from 3.6 to 18.2 mM, keeping the concentration of TBAC constant at 9.1 mM. The double logarithmic plot of the initial rate versus the concentration of the catalyst is shown in Figure 6a. Differently from the case of **5b**, the reaction order with respect to the complex **2** was found to be one. Moreover, the slope of the line does not change by varying the TBAC/**2** ratio, even in the case in which the co-catalyst is substoichiometric with respect to **2**. This is in agreement with the observation that the iron complex is able to catalyze the reaction even in the absence of an exogenous nucleophile.

An analogous set of experiments was performed to study the dependence with respect to TBAC, varying the concentration of TBAC in the range from 7.6 to 29.7 mM, keeping the concentration of **2** constant at 9.8 mM. The double logarithmic plot of the initial rate versus the concentration of the co-catalyst is shown in Figure 6b. A zero order dependence was found in this case. However, the reaction rate depends from the [TBAC] when the iron catalyst is in excess. That dependence indicates that the addition of TBAC results in a better activation of the complex, with the concentration of the activated catalyst increasing until the 1/1 molar ratio. The reaction order with respect to **5m** was determined performing a series of experiments with different concentration of the co-monomer.

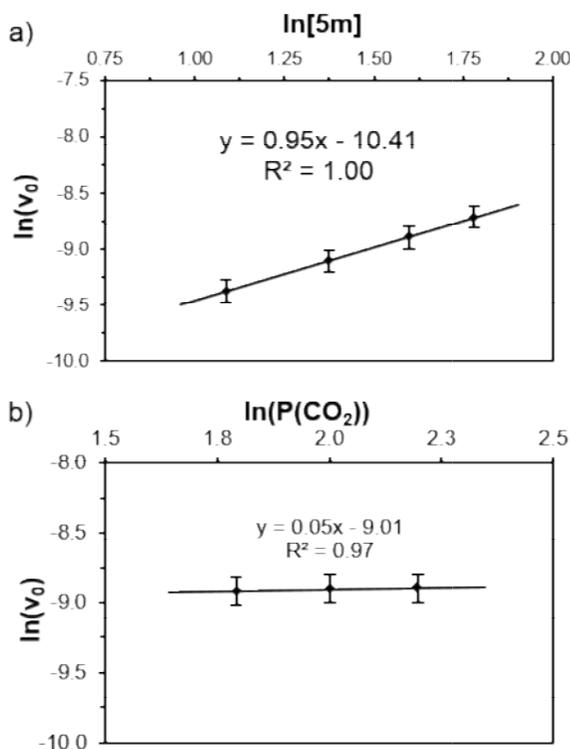


Figure 7. Plot of $\ln(v_0)$ for the formation of **6m** versus: (a) $\ln[5\mathbf{m}]$, (b) $\ln(P(\text{CO}_2))$.

The double logarithmic plot of the initial rate with respect of the concentration of the epoxide gives a first-order dependence

with respect to **5m** (Figure 7a). With constant **5m** concentration, the carbon dioxide pressure was changed in the range from 6 to 9 bar and the reaction order with respect to CO_2 was found to be zero (Figure 7b).

The co-polymerization reaction of **5m** with CO_2 was conducted, under the same reaction conditions, at different temperature in the range from 60 to 90 °C, and the Eyring plot is shown in Figure S42. Using this data we found the following results: $\Delta H^\ddagger = 11.9 \pm 0.3$ kcal/mol, $\Delta S^\ddagger = -36 \pm 2$ cal/(mol · K) and $\Delta G^\ddagger_{353} = 24.4 \pm 1.1$ kcal/mol.

Finally, the kinetic equation for the propagation step of the co-polymerization of **5m** with CO_2 can be written as follow:

$$v = k \cdot [5\mathbf{m}] \cdot [2] \quad (\text{eq. 4})$$

In analogy to the formation of **6b**, also in this case it is reasonable to assume that the rate limiting step involves the ring opening of the epoxide by the growing polymer chain. As a consequence, an excess of co-catalyst is not needed to obtain the maximum catalytic activity, differently from what observed in the case of the COCs formation. The reason lays in the fact that, in the case of the co-polymerization, the TBAC is involved only during the initiation step, i.e. the ring opening of the first epoxide ring. Indeed, during the propagation reaction, the insertion of the epoxide happens in the Fe-carbonate bond.

The divergence in the cyclic carbonate or the polycarbonate formation on the basis of the reaction conditions has been reported for different catalytic systems. In particular, the temperature dependence of the product mixture composition in the case of **5b** and **5m** was investigated by Coates,²⁸ Darensburg,^{18a} and Williams.^{13b} In all these cases, both the products were claimed to derive from the same reaction mechanism, in which the polycarbonate was identified as the kinetic product, favored at low temperature, and the cyclic carbonate as the thermodynamic product, favored at high temperature. The key for the selective formation of the polycarbonate at low temperature was thus identified in the different activation energies between the propagation and the backbiting reaction.

In the case of the [OSSO]-Fe(III) complexes **1-4** discussed in this study, a direct comparison of the activation energies for a given epoxide between the cyclic and polymeric product is not possible, due to the selective formation of only one product regardless the reaction conditions employed, underlying that the observed selectivity is rather a consequence of the complete inhibition of the cyclization for the epoxide **5m**.

Computational study. DFT calculations were performed to get insight into the mechanism for Fe catalyzed coupling of CO_2 with epoxides. The [OSSO]-Fe(III) complex **3** was selected as catalyst for the calculations. Considering that the aforementioned experimental studies showed that up to 2 equivalents of TBAB can react with the iron pre-catalyst (Figures 2-4), we started by exploring the substitution of MeCN of pre-catalyst **3** by a Br^- of TBAB, leading to the Fe-Br intermediate A (Figure 8) with an energy gain of 4.2 kcal/mol (Scheme S1). Based on this, we assumed complex A as the starting point of the catalytic cycle. The overall pathway (Figure 8) is divided into three sections: i) ring opening of **5b**; ii) CO_2 insertion; iii) ring closing to liberate the cyclic carbonate. Optimized geometries of selected transition states are shown in Figure 9.

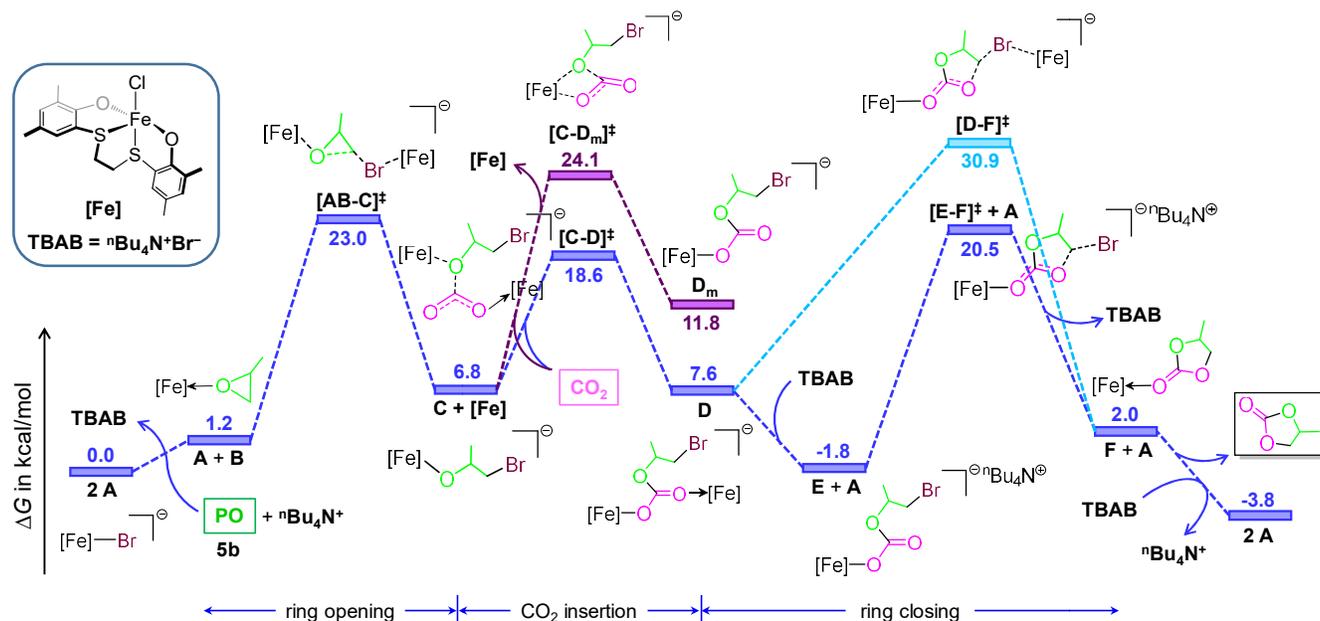


Figure 8. Computed energy profile for the cycloaddition of propylene oxide (**5b**) with CO₂ catalyzed by species **3**. For energy convention refer to the computational details.

The reaction starts with the replacement of Br[−] in **A** by the substrate (**5b**) generating **B**, a step endergonic by 1.2 kcal/mol. Based on the experimental evidence that the ring opening of **5b** is second order in [Fe] (Figure 2a), the most reasonable explanation consistent with the experiments is the interaction of **A** and **B** via the bimetallic transition state [AB-C][‡]. This step requires the overcoming of a free energy span of 23.0 kcal/mol from two **A**, and leads to the formation of the intermediate **C**, presenting a Fe–O σ-bond, and liberation of a Fe complex with a vacancy on the Fe center. The calculated energy span is in good agreement with the experimental activation energy, 19.0 ± 1.7 kcal/mol, estimated from the Eyring plot of Figure S35. Nevertheless, considering that the Fe-epoxide complex **B** is only 1.2 kcal/mol above **A**, and the large excess of **5b** relative to Fe and TBAB (Table 1 and Scheme 3), it is reasonable to assume the presence of a significant amount of **B** in solution.

Using a mixture of **A+B** as reference structure, the calculated energy barrier to transition state [AB-C][‡] is 21.8 kcal/mol, which is in better agreement with the experimental value, further supporting the mechanistic scenario we developed for the opening of **5b**.

The next step corresponds to insertion of CO₂ into the Fe–O bond of the intermediate **C** to generate the hemicarbonatate **D**. As no experiment is guiding the DFT calculations, we investigated both monometallic (purple) and bimetallic (blue) pathways. The bimetallic route,²⁹ via transition state [C-D][‡], is more facile due to a lower activation barrier from **C** (ΔG[‡] = 11.8 kcal/mol) compared to the monometallic route (ΔG[‡] = 17.3 kcal/mol).

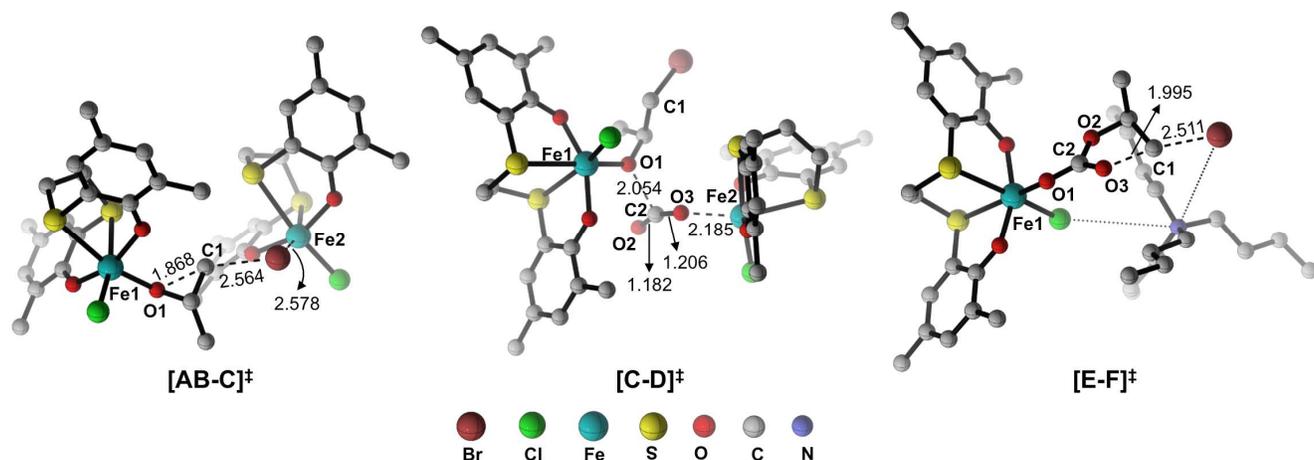


Figure 9. Optimized geometries of transition states (relative Gibbs free energy in kcal/mol). Hydrogen atoms are removed and phenyl groups of dppm ligands are shown in a simplified fashion for clarity.

This can be related to the greater electrophilic nature of the C atom of the Fe-activated CO₂ in transition state $[\mathbf{C-D}]^\ddagger$ ($q_{\text{C2}} = 1.092 e$), which facilitates the insertion of CO₂ with respect to the monometallic pathway, with a less electrophilic CO₂ in transition state $[\mathbf{C-D}_m]^\ddagger$ ($q_{\text{C2}} = 0.992 e$). The resulting hemicarbonates along the bimetallic route, \mathbf{D} , is also more stable (by 4.2 kcal/mol) than the congener \mathbf{D}_m along the monometallic route.

From intermediate \mathbf{D} we modeled the final ring closing step to achieve the desired cyclic carbonate, and we again tested different possibilities. The first was direct ring closing from intermediate \mathbf{D} . It occurs via transition state $[\mathbf{D-F}]^\ddagger$ and a total energy span of 30.9 kcal/mol from the reference two \mathbf{A} complexes. This energy span is too high to be consistent with the reaction condition used, so we searched for alternative pathways. We thus reacted \mathbf{D} with a TBAB molecule to release the neutral monomeric Fe complex \mathbf{E} with regeneration of \mathbf{A} . This is an exergonic step with release of 9.4 kcal/mol. Ring closing from \mathbf{E} via transition state $[\mathbf{E-F}]^\ddagger$, with an energy span of 22.3 kcal/mol, is instead consistent with the experimental conditions. In the resulting intermediate \mathbf{F} the formed cyclic carbonate is coordinated to Fe. Displacement of the cyclic carbonate by a TBAB molecule closes the catalytic cycle by regenerating the second starting species \mathbf{A} . The computed energy profile reveals that ring opening of $\mathbf{5b}$ is the rate determining step, and that TBAB and CO₂ are not involved in the rate controlling step. Finally, the experimental first order kinetics with respect to the $\mathbf{5b}$ concentration is also justified by the proposed mechanism.

The three fundamental transition states along the reaction pathway, $[\mathbf{AB-C}]^\ddagger$, $[\mathbf{C-D}]^\ddagger$ and $[\mathbf{E-F}]^\ddagger$, are shown in Figure 9. They exhibit breaking and forming bond distances comparable to those calculated for a bimetallic Fe(III) catalyst, supported by polydentate thioether-phenolate ligand.¹⁷ Nevertheless, for this bimetallic catalyst the CO₂ insertion step occurred on a single metal center, as the orientation of the breaking Fe-O bond on one site, and the forming Fe-CO₂ bond on the other site (nearly syn oriented) did not allow for an optimal interac-

tion between the molecular orbitals on the reacting atoms when two metal centers were involved.¹⁷ Differently, the monometallic catalysts of this work allow for an optimal orientation of the reacting groups, due to the almost anti orientation of the breaking Fe-O and forming Fe-CO₂ bonds in transition state $[\mathbf{C-D}]^\ddagger$ (Figure 9). On the other hand, the higher Fe-coordination environment in the bimetallic Fe-catalyst results in an easier ring-closing step,¹⁷ compared to the systems described in this work.

Having clarified the mechanism for the coupling of $\mathbf{5b}$ and CO₂ to a cyclic carbonate, we investigated the competition for cyclic carbonate formation versus co-polymer formation when $\mathbf{5m}$ is the epoxide. Based on the mechanism shown in Figure 8, \mathbf{E} is the key intermediate from which another epoxide will insert into the Fe-O bond within a co-polymerization process, or ring closing to cyclic carbonate will start. We have modeled these two competitive steps following the experimental conditions, *i.e.* in presence of TBAC with toluene as the solvent. Consistently with the experiments, the calculated energetics (Figure 10a) indicates that $\mathbf{5m}$ insertion into the Fe-O bond of \mathbf{E} , via transition state $[\mathbf{G}_{\text{CHO}}\text{-H}_{\text{CHO}}]^\ddagger$ (dark orange line), is favored by 4.9 kcal/mol over ring closing via transition state $[\mathbf{E}_{\text{CHO}}\text{-F}_{\text{CHO}}]^\ddagger$ (blue line).

For the sake of completeness, we also calculated the copolymerization step using $\mathbf{5b}$ as the epoxide in the presence of TBAC (Figure 10b). According to calculations, coordination and insertion of a $\mathbf{5b}$ molecule into the Fe-O bond of \mathbf{E} , via transition state $[\mathbf{G}_{\text{PO}}\text{-H}_{\text{PO}}]^\ddagger$, is disfavored by 9.0 kcal/mol relative to the ring-closing transition state $[\mathbf{E}_{\text{PO}}\text{-F}_{\text{PO}}]^\ddagger$. Consistently with experiments, ring-closure is thus favored with $\mathbf{5b}$, while co-polymerization is favored with $\mathbf{5m}$. Comparison of Figures 8 and 10b offers insight on the different role of the halide leaving group in the ring-closure step. The energy barrier for ring-closure with Br⁻, 22.3 kcal/mol from \mathbf{E} , is calculated to be 5.4 kcal/mol higher than that calculated with Cl⁻, 16.9 kcal/mol. The higher barrier with Br⁻ can be probably related to poorer electrostatic stabilization of the larger bromide anion by the Bu₄N⁺ cation and by the solvent.

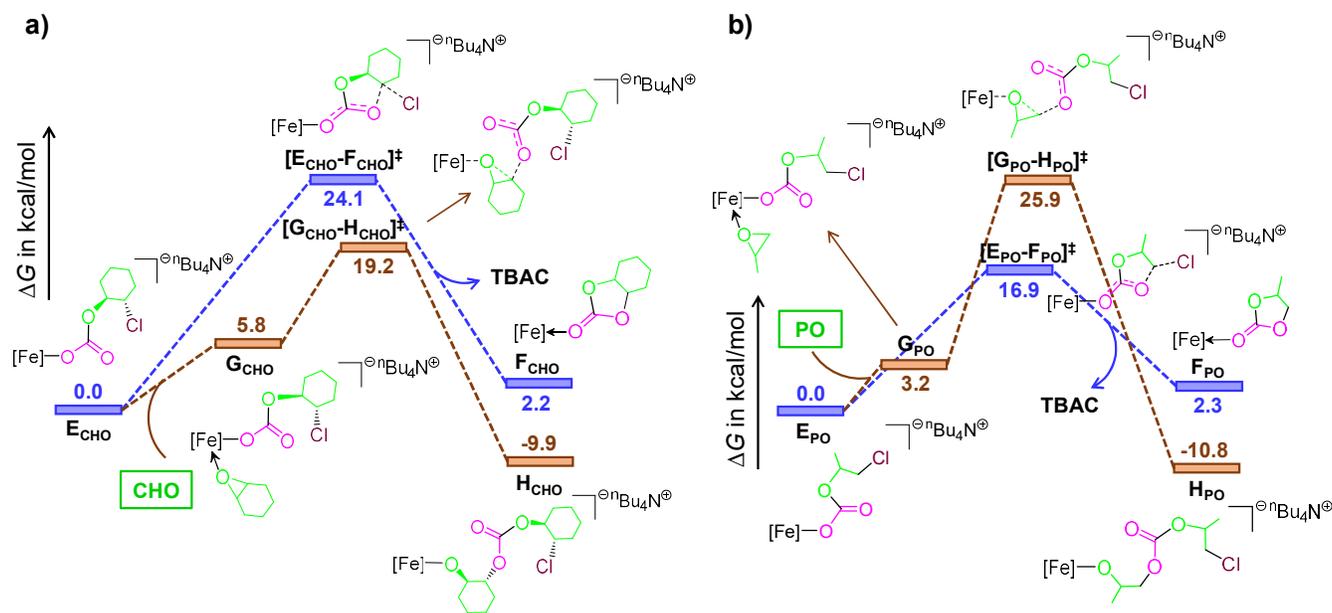


Figure 10. Energy profile for the competitive ring closing vs. copolymerization steps with $\mathbf{5m}$ (a) and $\mathbf{5b}$ (b). For energy convention refer to the computational methods.

To understand the different tendency of **5m** and **5b** to undergo co-polymerization versus cyclic carbonate formation, we compared the thermochemistry of cyclic carbonates formation versus the epoxides structure (Figure 11). For better accuracy, thermochemistry was calculated according to the protocol described in ref. 30 (see the SI, Scheme S2).

According to calculations, the conversion of **5m** into the corresponding cyclic carbonate is 3.8 kcal/mol harder than conversion of **5b**, and 3.0 kcal/mol harder than conversion of **5j**. This trend is well in line with the tendency of **5b** and **5j** to undergo cyclic carbonate formation, and the tendency of **5m** to undergo co-polymerization.

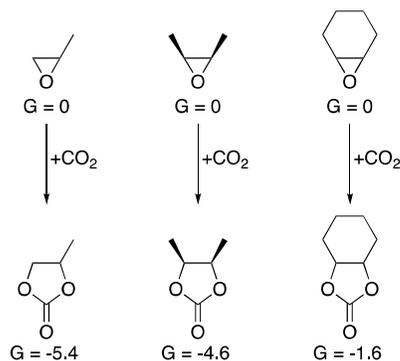


Figure 11. Thermochemistry of cyclic carbonate formation, in kcal/mol.

CONCLUSION

In conclusion, here we have reported on a new family of Fe(III) complexes (**1-4**) supported by bis-thioether-diphenolate [OSSO]-type ligands and their ability to promote the reaction of CO₂ with oxiranes under very mild reaction conditions. These catalysts display excellent activity and selectivity for the formation of terminal and internal COCs starting from the corresponding epoxides. Intriguingly, in the case of cyclohexene oxide, the formation of the corresponding polycarbonate without the concomitant production of the cyclic carbonate, regardless the reaction temperature, was observed.

Moreover, kinetic measurements revealed a different reaction order for the iron complex in the cycloaddition and in the co-polymerization reactions. As matter of fact, in the case of the formation of propylene carbonate, a second order with respect to the Fe(III) concentration was found, while in the case of the poly(cyclohexene carbonate) a first order was observed, even though the ring opening of the epoxy-ring was found to be the rate limiting step in both cases. DFT based calculations identified a key monometallic Fe-carbonate intermediate from which both cyclization and co-polymerization can start. However, only in the case of cyclohexene oxide the energy barrier of the chain propagation step was found to be lower than that of the ring closing of the cyclic product.

These results show the potential of iron based catalytic systems for these reactions under low CO₂ pressure. Synergic experimental and computational investigation provided a detailed understanding of the overall mechanism, including the factors governing the selectivity in the formation of the cyclic versus the polymeric product. This knowledge can pave the way for the design of highly chemoselective catalysts, based on Earth-crust abundant metals such iron.

ASSOCIATED CONTENT

Supporting Information. Details on: synthesis and characterization of complexes, crystal structure, reaction procedures, ATR FT-IR kinetic measurements and DFT calculations are supplied as Supporting Information. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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Column 1	Column 2	Column 3	Column 4	Column 5	Column 6	Column 7	Column 8

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